

Surface Chemistry Ncert Solutions

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Solution: When a gas is adsorbed on the surface of a solid, its entropy decreases, i.e., ΔS is negative. Now, $\Delta G = \Delta H - T\Delta S$
For a process to be spontaneous, ΔG must be negative. As here, ΔS is negative, therefore, $T\Delta S$ is positive ΔG can be negative only if ΔH is negative
Hence, adsorption is always exothermic.
5.9.

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we come across in surface chemistry may be pure compounds or solutions. The interface is normally a few molecules thick but its area depends on the size of the particles of bulk phases. Many important phenomena, noticeable amongst these being corrosion, electrode processes, heterogeneous catalysis, dissolution and crystallisation occur at interfaces. The subject of surface chemistry finds

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Surface Chemistry - NCERT

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Chemistry Notes for class 12 Chapter 5 Surface Chemistry Surface Chemistry is the branch of chemistry which deals with the phenomenon that occurs on the surfaces or interfaces, such phenomenon includes corrosion. catalysis, crystallisation, etc

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Important Questions for CBSE Class 12 Chemistry - Surface Chemistry Previous Years' Questions 2015 Very Short Answer Type Questions [1 Mark] 1. Out of AlCl_3 and NaCl , which is more effective in causing coagulation of a negative sol and why? Answer: AlCl_3 , because Al^{3+} has higher charge than Na^+ ion. Higher the charge, more effective [...]

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